Name KEY Date _____ Class 1ST Qtrly

3. How many neutrons are in each atom?
   a. $^{23}_{11}Na$ 12
   c. $^{81}_{35}Br$ 46
   b. $^{238}_{92}U$ 146
   d. $^{19}_{9}F$ 10

4. The two most abundant isotopes of carbon are carbon-12 (mass = 12.00 amu) and carbon-13 (mass = 13.00 amu). Their relative abundances are 98.9% and 1.10%, respectively. Calculate the atomic mass of carbon. 12.011 amu

5. Element X has two isotopes: X-100 and X-104. If the atomic mass of X is 101 amu, what is the relative abundance of each isotope in nature?
   75% X-100
   25% X-104