

KEY

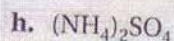
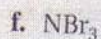
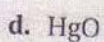
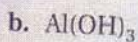
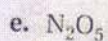
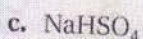
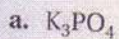
Date _____

Class _____

2. Complete this table by writing correct formulas for the compounds formed by combining positive and negative ions.

	SO_4^{2-}	NO_3^-	OH^-	PO_4^{3-}
Ca^{2+}	CaSO_4	$\text{Ca}(\text{NO}_3)_2$	$\text{Ca}(\text{OH})_2$	$\text{Ca}_3(\text{PO}_4)_2$
Al^{3+}	$\text{Al}_2(\text{SO}_4)_3$	$\text{Al}(\text{NO}_3)_3$	$\text{Al}(\text{OH})_3$	AlPO_4
Na^+	Na_2SO_4	NaNO_3	NaOH	Na_3PO_4
Pb^{4+}	$\text{Pb}(\text{SO}_4)_2$	$\text{Pb}(\text{NO}_3)_4$	$\text{Pb}(\text{OH})_4$	$\text{Pb}_3(\text{PO}_4)_4$

3. Name the following compounds.



4. Explain the difference between the law of definite proportions and the law of multiple proportions.

D- Elements in same proportions

M- Masses of elements in 2 compounds with same 2 elements will be in small whole # ratios